

GCE

Chemistry B

Unit **H433A/02**: Scientific literacy in chemistry

Advanced GCE

Mark Scheme for June 2017

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This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which marks were awarded by examiners. It does not indicate the details of the discussions which took place at an examiners' meeting before marking commenced.















All examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes should be read in conjunction with the published question papers and the report on the examination.

OCR will not enter into any discussion or correspondence in connection with this mark scheme.















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Annotations available in RM Assessor

Annotation	Meaning
	Correct response
	Incorrect response
	Omission mark
	Benefit of doubt given
	Contradiction
	Rounding error
	Error in number of significant figures
	Error carried forward
	Level 1
	Level 2
	Level 3
	Benefit of doubt not given
	Noted but no credit given
	Ignore

12. Abbreviations, annotations and conventions used in the detailed Mark Scheme (to include abbreviations and subject-specific conventions).

Annotation	Meaning
/	alternative and acceptable answers for the same marking point
✓	Separates marking points
DO NOT ALLOW	Answers which are not worthy of credit
IGNORE	Statements which are irrelevant
ALLOW	Answers that can be accepted
()	Words which are not essential to gain credit
—	Underlined words must be present in answer to score a mark
ECF	Error carried forward
AW	Alternative wording
ORA	Or reverse argument

Annotation	Meaning
	Correct response
	Incorrect response
	Calculation mark
	Result of division given
	Conradiction
	Rounding error
	Error in number of significant figures
	Error carried forward
	Level 1
	Level 2
	Level 3
	Result of division not given
	Result but no result given
	Ignore

Subject-specific Marking Instructions**INTRODUCTION**

Your first task as an Examiner is to become thoroughly familiar with the material on which the examination depends. This material includes:

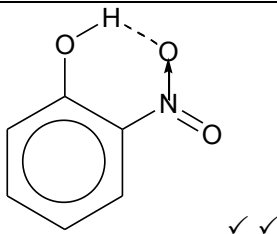
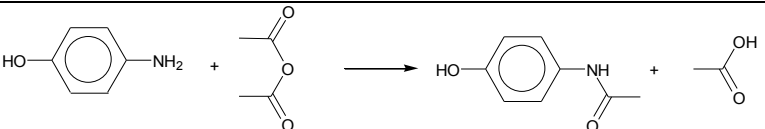
- the specification, especially the assessment objectives
- the question paper
- the mark scheme.

You should ensure that you have copies of these materials.

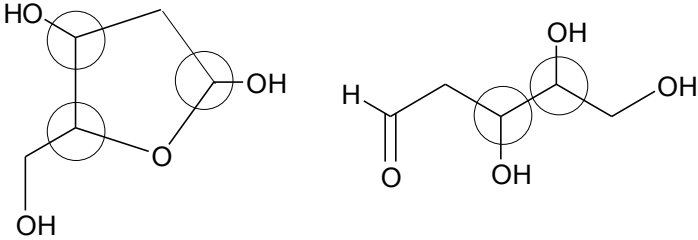
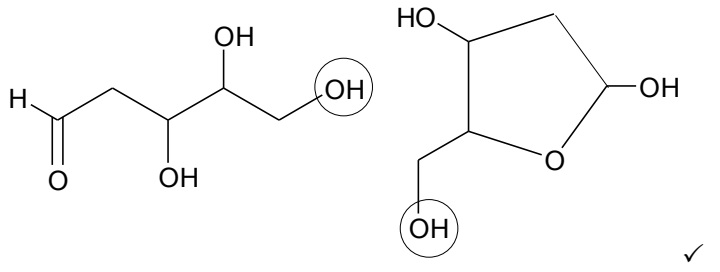
You should ensure also that you are familiar with the administrative procedures related to the marking process. These are set out in the OCR booklet **Instructions for Examiners**. If you are examining for the first time, please read carefully **Appendix 5 Introduction to Script Marking: Notes for New Examiners**.

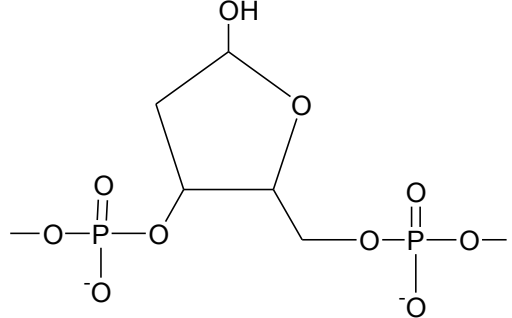
Question			Answer	Marks	Guidance
1	(a)	(i)	$2\text{CH}_3\text{COOH} + \text{Na}_2\text{CO}_3 \rightarrow 2\text{CH}_3\text{COONa} + \text{CO}_2 + \text{H}_2\text{O}$ formulae ✓ balancing of correct formulae ✓	2	ALLOW any unambiguous formulae (including molecular formulae) ALLOW H_2CO_3 as a product in a balanced equation for 1 mark IGNORE state symbols
		(ii)	FIRST CHECK ANSWER LINE If answer = 11.25 or 11.3 (cm³) award 2 marks amount $\text{CH}_3\text{COOH} = 25 \times 0.450/1000$ OR $0.01125(\text{mol})$ ✓ volume $\text{Na}_2\text{CO}_3 (= 0.5 \times 0.01125 \times 1000/0.500)$ = 11.25 (cm ³) ✓	2	ALLOW ecf from 1:1 ratio in a(i) ALLOW 3 or more sf
	(b)		$\text{CH}_3\text{COO}^- + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COOH} + \text{OH}^-$	1	IGNORE state symbols ALLOW arrow for equilibrium sign
	(c)	(i)	$\text{CH}_3\text{COOH} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}^+$	1	ALLOW: $\text{CH}_3\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}_3\text{O}^+$ Equilibrium sign required IGNORE state symbols
		(ii)	FIRST CHECK ANSWER LINE If answer = 3.1(462) award 2 marks $[\text{H}^+] = \sqrt{(1.7 \times 10^{-5} \times 0.030)}$ OR $7.14(\dots) \times 10^{-4}$ ✓ $\text{pH} (= -\log 7.14 \times 10^{-4}) = 3.1(462)$ ✓	2	ALLOW 'H ⁺ ' for '[H ⁺]' ALLOW ecf for second mark provided value for [H ⁺] is quoted and it is smaller than 3×10^{-2} and greater than 1.1×10^{-7} .
	(d)		acid/ H^+ moves equilibrium to left / reactants ✓ idea of restoring/maintaining pH ✓ large concentrations/ amounts/excess of salt/ CH_3COO^- ✓	3	Equilibrium must be written out (either for ethanoic acid or HA) to score first mark (or they could refer back to (c)(i)) ALLOW idea of "ethanoate ions react with H^+ ions to restore equilibrium" to score first mark
	(e)	(i)	$([\text{H}^+] = K_a \times [\text{CH}_3\text{COOH}]/[\text{CH}_3\text{COO}^-])$ gives $\text{pH} = 4.77/4.8$ ✓	1	

Question			Answer	Marks	Guidance
		(ii)	FIRST CHECK ANSWER LINE If answer = 0.35(g) award 4 marks $[\text{CH}_3\text{COO}^-] = K_a \times [\text{CH}_3\text{COOH}] / [\text{H}^+] \checkmark$ $= 1.7 \times 10^{-5} \times 0.1 / 10^{-5} \text{ OR } 0.17 \text{ (mol dm}^{-3}\text{)} \checkmark$ mass CH_3COONa per $\text{dm}^3 = 82 \times 0.17 \text{ OR } 13.94 \text{ g } \checkmark$ mass per $25 \text{ cm}^3 = 13.94/40 = 0.35 \text{ g } \checkmark$ OR moles in $25\text{cm}^3 = 0.17/40 = 4.25 \times 10^{-3} \checkmark$ mass per $25 \text{ cm}^3 = 4.25 \times 10^{-3} \times 82 = 0.35\text{g } \checkmark$	4	Accept 0.349g / 0.3485g ALLOW 2 or more sf ALLOW ecf throughout ALLOW correct expression for K_a
			Total	16	

Question			Answer	Marks	Guidance
2	(a)		Phenol/hydroxy(l) ✓ <u>secondary</u> amide ✓	2	NOT alcohol
	(b)	(i)	FIRST CHECK ANSWER LINE If answer = 13.5 / 14(g) award 2 marks amount 4-nitrophenol = 5.0/139 OR 0.03597 (mol) AND mass phenol to give 100% = 5.0 x 94/139 OR 3.381 (g) ✓ scaling by 100/25 to get 13.5 / 14 (g) ✓	2	ALLOW ecf Any number scaled by 100/25 and to 2 or 3 sf scores second mark (if first mark not scored)
		(ii)		2	ALLOW on NO ₂ group: <ul style="list-style-type: none"> • double bond to either oxygen with a single or dative bond to the other • 'one and a half' bonds to each oxygen One mark for correct bonding within NO ₂ group One mark for hydrogen bond between correct H and O (even if bonding wrong)
		(iii)	reduction AND amine	1	
		(iv)	 reactants ✓ products ✓	2	IGNORE non-skeletal formulae /ambiguous attachments Allow correct use of Ethanoyl chloride for 1 mark
		(v)	dissolve in minimum volume of <u>hot</u> water / solvent ✓ filter (hot solution) removing insoluble impurities ✓ allow to crystallise /AW✓ filter, soluble impurities removed/remain in solution ✓	4	ALLOW wash (and dry) crystals - soluble impurities are washed away ✓

Question			Answer	Marks	Guidance
		(c)(i)	(AM404) has a similar shape to andanamide AW ✓ (AM404) fits/ binds/bonds to active site ✓ (AM404 in active site) not broken down/ stays on (active site)/ blocks site to/ competes with andanamide AW ✓	3	
		(ii)	(all) cis/Z ✓	1	
		(iii)	lack of rotation/twisting ✓	1	
			Total	18	

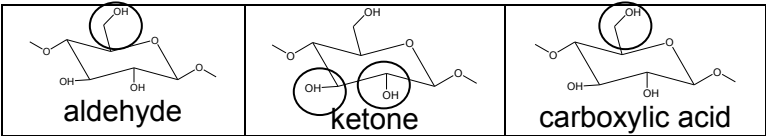
Question			Answer	Marks	Guidance								
3	(a)	(i)		1	All 5 Carbon atoms must be circled								
		(ii)	aldehyde ✓	1	NOT carbonyl here								
		(iii)	<div><div>one from</div><table><tr><td>test 1 mark ✓</td><td>result linked to appropriate test 1 mark ✓</td></tr><tr><td>heat with Fehling's/ Benedicts solution</td><td>(brick) red ppt</td></tr><tr><td>add Tollens' reagent and warm</td><td>silver mirror /AW</td></tr><tr><td>heat with acid dichromate</td><td>goes green</td></tr></table></div>	test 1 mark ✓	result linked to appropriate test 1 mark ✓	heat with Fehling's/ Benedicts solution	(brick) red ppt	add Tollens' reagent and warm	silver mirror /AW	heat with acid dichromate	goes green	2	Reagents may be specified (eg 'silver nitrate and ammonia' for Tollens')
test 1 mark ✓	result linked to appropriate test 1 mark ✓												
heat with Fehling's/ Benedicts solution	(brick) red ppt												
add Tollens' reagent and warm	silver mirror /AW												
heat with acid dichromate	goes green												
		(iv)	 <p>OH/it is attached to a C with 2H/ one C-C bond /attached to a carbon that is only bonded to one other carbon / R group✓</p>	2	Both OH groups must be circled								

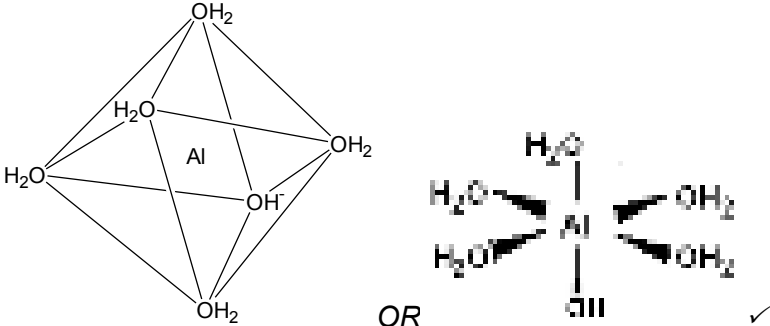
Question			Answer	Marks	Guidance
	(v)		Water / a small molecule is not produced/ both have same molecular formula	1	
	(vi)		$-\text{CHO} + \text{HO}- \rightarrow -\text{CH}(\text{OH})-\text{O}-$	1	ALLOW any unambiguous representation
(b)			 <p>correct bonding in phosphates ✓ correct connection of phosphates AND correct ring structure ✓</p>	2	
(c)	(i)		Base(s) ✓ Condense/react with(remaining) OH on deoxyribose ✓	2	
	(ii)		GAC	1	
	(iii)		leucine ✓	1	
	(iv)		mRNA sequence /(triplet) codon ✓ codes for/fits with tRNA (anti-codon) on amino acid ✓	2	

	(d)	<p>Please refer to the marking instructions on page 4 of this mark scheme for guidance on how to mark this question.</p> <p>Level 3 (5-6 marks) Structure correct and at least one piece of evidence related to the structure is provided from each spectrum.</p> <p><i>There is a well-developed line of reasoning which is clear and logically structured. The information presented is relevant and substantiated.</i></p> <p>Level 2 (3-4 marks) Structure correct but does not use evidence from each spectrum OR structure not given or incorrect but at least four correct pieces of evidence given from a minimum of two spectra</p> <p><i>There is a line of reasoning presented with some structure. The information presented is relevant and supported by some evidence.</i></p> <p>Level 1 (1-2 marks) Structure correct with no evidence OR Structure not given or incorrect but at least two correct pieces of evidence given</p> <p><i>There is an attempt at a logical structure with a line of reasoning. The information is in the most part relevant.</i></p> <p>0 marks <i>No response or no response worthy of credit.</i></p>	6	<p>indicative scientific points include:</p> <p>structure $\text{CH}_3\text{OOCCH}(\text{OCH}_3)_2$ Allow $(\text{CH}_3\text{O})_2\text{CHOCOCH}_3$</p> <p>infrared spectrum</p> <ul style="list-style-type: none"> • C=O for ester (at 1750cm^{-1}) • C—O for ester/ether (at $1000 - 1300\text{cm}^{-1}$) • no O—H/COOH (at $2500 - 3000\text{cm}^{-1}$) <p>ignore idea of C—H stretch/O—H stretch at approx. 3000cm^{-1}</p> <p>^1H NMR</p> <ul style="list-style-type: none"> • 3 proton environments • no splitting • 2 CH_3 groups in the same environment • all O—CH (AW) <p>^{13}C NMR</p> <ul style="list-style-type: none"> • four C environments • C=O at $\delta = 160\text{ppm}$ • C—O at $\delta = 50\text{ppm}$ • no C—C present <p>If incorrect bond or environment identified consider whether or not the line of reasoning has been impeded and if so, then award lower mark within a level</p>
		Total	22	

Question			Answer				Marks	Guidance
4	(a)	(i)	Equation no.	Oxidation state of Cr in reactant	Oxidation state of Cr in product	Has Cr been oxidised, reduced or neither?		4 Mark each row separately Penalise (+) omitted /3+ etc once only
			4.1	+3	+6	oxidised	✓	
			4.2	+6	+6	neither	✓	
			4.3	+6	+3	reduced	✓	
			4.4	+3	0	reduced	✓	
		(ii)	FIRST CHECK ANSWER LINE If answer = 0.46 or rounding to 0.465 (kg) award 2 marks amount chromite = 1000/223.8 OR 4.468 (mol) ✓ mass Cr = 2 x 4.468 x 52/1000 = 0.46467 (kg) ✓				2	ALLOW 2 or more sf ALLOW 0.232kg for 1 mark
	(b)	(i)	$1s^2 2s^2 2p^6 3s^2 3p^6 3d^3$				1	ALLOW any sized letters but numbers must be superscripts Accept [Ar] $3d^3$ ALLOW $4s^0$
		(ii)	beaker (etc) with solution and two electrodes with battery/power supply connected ✓ solution labelled chromium(III) chloride / CrCl_3 / Cr^{3+} ✓ +ve electrode/anode labelled 'graphite' AND -ve electrode/cathode labelled 'steel (object)'. ✓				3	If two beaker diagram drawn, allow 1 mark for a steel electrode inserted into a chromium(III)chloride solution Ignore battery convention UNLESS polarity of electrodes not otherwise indicated
		(iii)	$\text{Cr}^{3+} + 3e^{(-)} \rightarrow \text{Cr}$ ✓				1	IGNORE state symbols ALLOW ecf from labelled cathode half cell if present in b(ii)

		(iv)	FIRST CHECK ANSWER LINE If answer = 8 (hours) award 3 marks moles of electrons = $3 \times 26/52$ OR 1.5 ✓ time = $1.5 \times 96500/5$ OR 28950 (sec) ✓ time in hours = $28950/3600 = 8(.04)$ (hours) ✓	3	ALLOW ecf ALLOW any sf If final answer rounds to 2.7 hours scores 2
	(c)	(i)	<u>chloride</u> (ions) AND water (molecules)	1	IGNORE formulae NOT <u>chlorine</u> ALLOW chloro and aqua
		(ii)	Add a named ionic chloride (solution) or hydrochloric acid/HCl ✓ to move equilibrium to the left/reactants ✓	2	Mark independently
	(d)	(i)	FIRST CHECK ANSWER LINE If answer = 0.977 (g/100cm³) on second answer line award 6 marks If answer = 0.21 (mol dm⁻³) on first answer line award 5 marks amount $\text{Cr}_2\text{O}_7^{2-}$ init. = $20 \times 0.2/1000$ OR 4×10^{-3} (mol) ✓ amount $\text{Na}_2\text{S}_2\text{O}_3 = 27.6 \times 0.1/1000$ OR 2.76×10^{-3} (mol) ✓ amount $\text{Cr}_2\text{O}_7^{2-}$ left = $2.76 \times 10^{-3}/6$ OR 4.6×10^{-4} (mol) ✓ amount $\text{Cr}_2\text{O}_7^{2-}$ used = 3.54×10^{-3} (mol) ✓ conc ethanol = $(1.5 \times 3.54 \times 10^{-3} \times 40)$ = $0.212(4)$ (mol dm ⁻³) ✓ % (= 0.0212×46) = $0.975/0.977$ (g/100cm ³) ✓	6	ALLOW two or more sf ALLOW ecf throughout.
	(d)	(ii)	no other oxidising agents/ reducing agents in the beer	1	
			Total	24	

Question	Answer	Marks	Guidance
5 (a)	 <p>aldehyde ketone carboxylic acid</p>	3	<p>ALLOW just OH circled or C as well (as shown) Both circles must be shown for the ketones. If only carbon atoms have been identified, mark incorrect once and apply ecf</p>
(b) (i)	<p><i>Please refer to the marking instructions on page 4 of this mark scheme for guidance on how to mark this question.</i></p> <p>Level 3 (5 – 6 marks) Shows a good understanding of all three areas</p> <p><i>There is a well-developed line of reasoning which is clear and logically structured. The information presented is relevant and substantiated.</i></p> <p>Level 2 (3 – 4 marks) Shows a sound understanding of at least two areas. OR Discusses all three areas, demonstrating a sound understanding of one area</p> <p><i>There is a line of reasoning presented with some structure. The information presented is relevant and supported by some evidence.</i></p> <p>Level 1 (1 – 2 Marks) Shows some understanding of one area.</p> <p><i>There is an attempt at a logical structure with a line of reasoning. The information is in the most part relevant.</i></p> <p>Level 0 (0 marks) No response or response has no merit.</p>	6	<p>Indicative scientific points include</p> <p>Why radiation absorbed</p> <ul style="list-style-type: none"> • electrons move to higher energy levels/shells • absorbing light/radiation • frequency absorbed $\Delta E = h\nu$ (allow $E = h\nu$ if clear reference to energy gap) <p>Affect of structure on frequency of radiation</p> <ul style="list-style-type: none"> • amount of delocalisation affects ΔE/ frequency • smaller delocalisation, larger ΔE/frequency • lignin has smaller chromophore/ less delocalisation than decomposition prods <p>Source of yellow colour</p> <ul style="list-style-type: none"> • uv higher frequency/ larger ΔE than visible • look yellow because they absorb the complementary colour/ blue/violet <p>For answers that talk about electrons falling and releasing radiation/ light/colour impedes the line of reasoning and should result in the lower mark within a level being awarded.</p>

Question	Answer	Marks	Guidance
(ii)	<p>FIRST CHECK ANSWER LINE If answer = 342 (kJ mol⁻¹) award 3 marks</p> <p>Rearrangement of $E = h\nu$ and $c = \nu\lambda$ to $E = hc/\lambda$ ✓</p> <p>$E = 6.63 \times 10^{-34} \times 3 \times 10^8 / 3.5 \times 10^{-7}$ (J per atom) OR $E = 6.63 \times 10^{-34} \times 8.57 \times 10^{14}$ ✓</p> <p>multiply by N_A and divide by 1000 and evaluate $(6.63 \times 10^{-34} \times 3 \times 10^8 \times 6.02 \times 10^{23} / 3.5 \times 10^{-7} \times 1000)$ $= 342 \text{ kJ mol}^{-1}$ ✓</p>	3	<p>ALLOW ecf MP2 subsumes MP1 and scores 2 marks</p>
(c) (i)	<p>$[\text{Al}(\text{H}_2\text{O})_6]^{3+}$ / H_2O <u>ligand/ in complex</u> ✓</p>	1	<p>Allow water/aqua for H_2O.</p>
(ii)	 <p>octahedral ✓</p>	2	<p>IGNORE charges</p> <p>Allow diagrams that are unambiguous in showing adjacent equatorial ligands with two coming out of the plane, and two going into the plane of the paper</p> <p>Do Not allow bonds to H atoms, must be to O as bonding is to the lone pair of electrons</p> <p>Mark independently</p>

	(d)	$2\text{Fe}^{2+} + \text{O}_2 + 2\text{H}^+ \rightarrow 2\text{Fe}^{3+} + \text{H}_2\text{O}_2$ ✓ Idea that Fe^{2+} is not recycled ✓	2	
	(e)	Use of diethyl zinc/ $(\text{C}_2\text{H}_5)_2\text{Zn}$ ✓ Adding nanoparticles/ microparticles of $\text{MgO}/\text{Mg}(\text{OH})_2/\text{Ca}(\text{OH})_2/\text{metal hydroxides}$ ✓ $(\text{C}_2\text{H}_5)_2\text{Zn} + 2\text{H}^+ \rightarrow \text{Zn}^{2+} + 2\text{C}_2\text{H}_6$ OR $\text{Ca}(\text{OH})_2 + 2\text{H}^+ \rightarrow \text{Ca}^{2+} + \text{H}_2\text{O}$ OR $\text{Mg}(\text{OH})_2 + 2\text{H}^+ \rightarrow \text{Mg}^{2+} + \text{H}_2\text{O}$ ✓	3	
		<i>Total</i>	20	

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