

Name: _____

1. The atomic number of an atom is _____
_____ (2)

2. The 3 subatomic particles (particles in an atom) are,
(i) _____
(ii) _____
(iii) _____ (6)

3. In an atom the number of protons equals the number of electrons.
Is this True or False? Answer: _____ (2)

4. Atoms of the same element, e.g. Carbon, but with different numbers of
neutrons in the nucleus are called _____
(2)

5. In the boxes provided draw the Bohr Diagram of each of the following,
Give the number of Protons and Neutrons in the centre and draw the electrons
in orbits around the nucleus.

(i) Carbon

(ii) Sodium



(6)

6. Give the electron configurations for the two elements above.
Carbon = (__, __) Sodium = (__, __, __) (4)

7. If an element has 2 electrons in its outer shell, which Group will it belong to?
Group Number = _____ Group Name=_____ (4)
8. Lithium and Sodium are all members of Group 1. What's the name of this group? Answer: _____ (2)
9. If an element has 2 electron shells which period does it belong to?
Answer: n = ____ (2)
10. List 2 metals and 2 non-metals from the periodic table
Metals = (i) _____ (ii) _____
Non-Metals = (i) _____ (ii) _____ (4)
11. Put the symbol **+**, **-** or **No Charge** next to the following particles.

Proton	
Electron	
Neutron	

(6)

